

CHM 1045 Chapter 2 Homework

1. Review Examples 2.03, 2.04, and 2.05. Write formulas and systematic names for the following compounds. You can use [this table](#), [this figure](#), and [this table](#). You can find the Latin ion names [here](#). Do not use prefixes for ionic compounds. (2 pts)

ferrous sulfate

ferric sulfate

cuprous fluoride

cupric fluoride

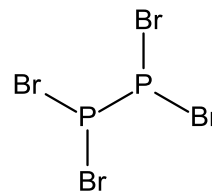
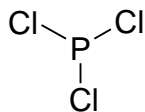
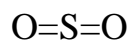
2. Write the systematic (Stock) and the older ([common / Latin](#)) names for the following compounds. You can use [this table](#), [this figure](#), and [this table](#). (2 pts)

CuCrO_4

$\text{Co}(\text{MnO}_4)_2$

3. Review the chapter notes and describe what a hydrate is. Include the overall phase of the substance. Write the balanced chemical reaction equation for anhydrous iron(II) sulfate becoming a heptahydrate, where “hepta-“ means 7. Include the subscripts, such as s, L, g, and aq, for physical phases of reactants and products. Name the product also. Refer to hydrates, chemical reactions, and Example 2.10 on the last page of the chapter 2 notes. (2 pts)

4. Review Examples 2.06, 2.07, and 2.08, as well as the sections on binary compounds in the text and chapter notes. Write the molecular formulas and names for the following binary molecules. Use prefixes and put the atoms in the proper order. (2 pts)



5. Write the systematic names for Na_2SO_4 and $(\text{NH}_4)_2\text{S}_2\text{O}_3$.
Write the systematic and common (bi) names for NaHCO_3 .
Do not use prefixes for ionic compounds.
You can use [this table](#), [this figure](#), and [this table](#). (2 pts)